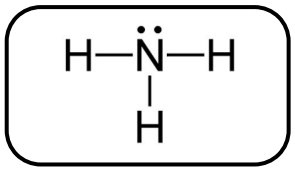
POLARITY

chapter 12 & 13



_____ electron distribution

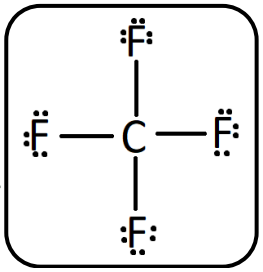
definition: the ability of atoms _____ shared electrons

difference in electronegativity: about 1.9 -0.5

need two (same/different) elements

electrons are hogged, leading to partial negatives or positives

_____ electron distribution



difference in electronegativity: about 0.4-0.0

bonds can be polar, but a _____ shape cancels out dipoles

no partial charges, electrons are shared _____

need two (same/different) elements

electrons are _____

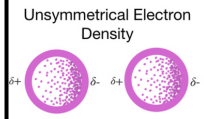
difference in electronegativity: 2.0 and greater.

very strong bond

Types of Forces

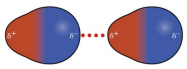
all molecules; _____ attractive forces due to movement of electrons.

ex. Br2



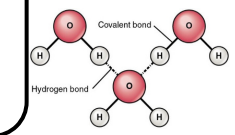
two polar molecules; the positive end of one molecule is attracted to the negative end

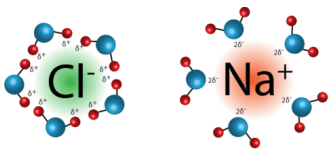
ex. HBr



attractive forces between polar molecules of H and _____

ex. H2O



When ionic compounds dissociate in a polar molecule

ex. NaCl